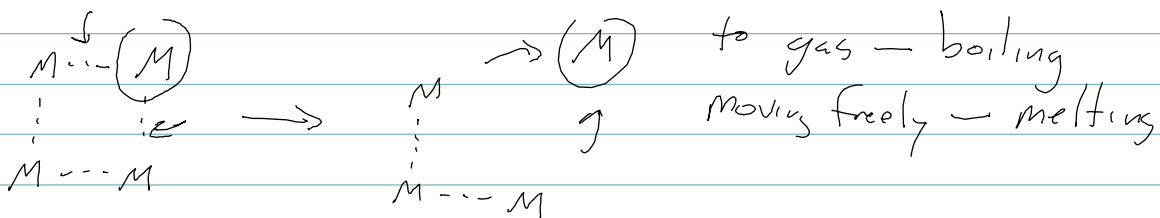


Boiling & Melting Point Trends due to Intermolecular Interactions



Melting & boiling require breaking intermolecular bonds.

Stronger I.M. / non-covalent bonds →
more energy to break →
higher b.p. or m.p.

Higher b.p. or m.p. { ion-ion > ion-dipole > H-bonds > dispersion only. } Lower b.p. or m.p.

b.p. { ↑ size, ↑ mass } ↓ size, ↓ mass

m.p. { densely packed molecules } loosely packed molecules.