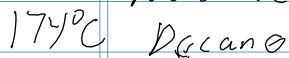
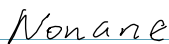
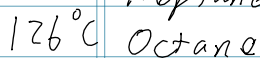
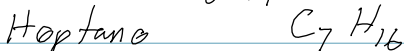
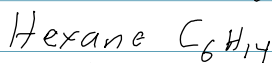
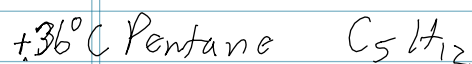
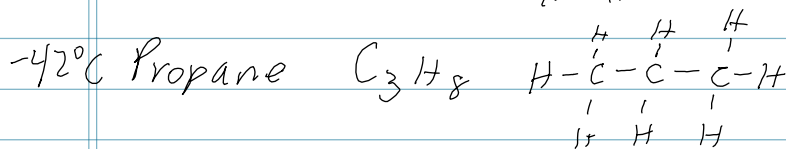
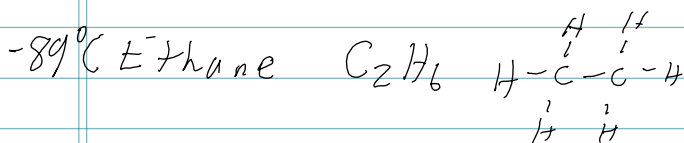
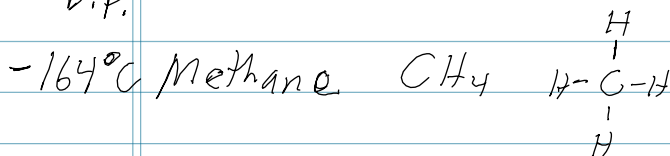


Alkanes

Hydrocarbon - contains only C & H

- ↳ Alkane - only single bonds
 - ↳ Linear: all C-C bonds in single chain
 - ↳ Branched (p2)
 - ↳ Cyclic (p3)

b.p.



natural gas

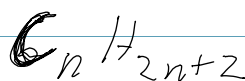
liquefied fuels
(under pressure)

109.5°

non-polar

$$\Delta EN = 0.3$$

$$C = 2.5 \quad H = 2.2$$



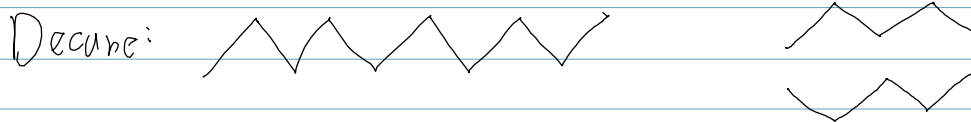
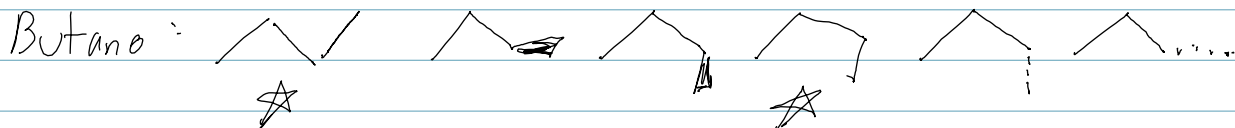
liquids → dispersion forces
low boiling points

gasoline: C_5H_{12} to $C_{12}H_{26}$

heating oil > $C_{12}H_{26}$

Conformational Flexibility

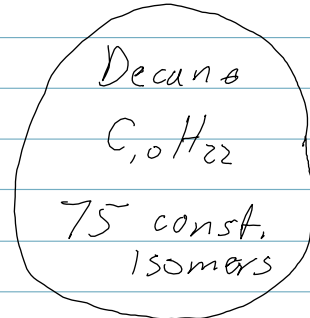
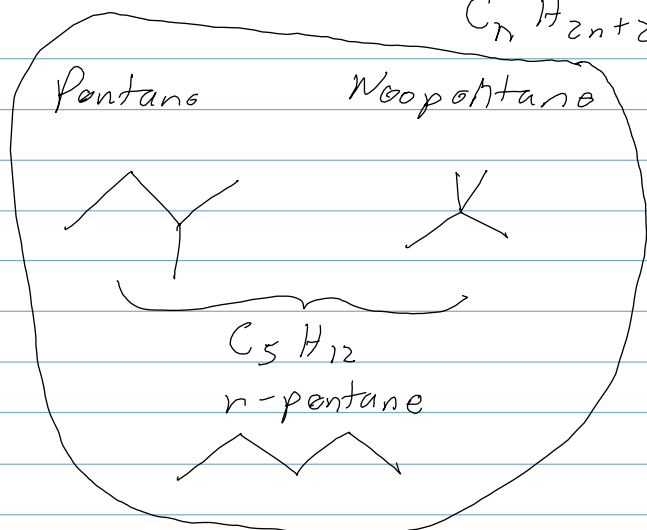
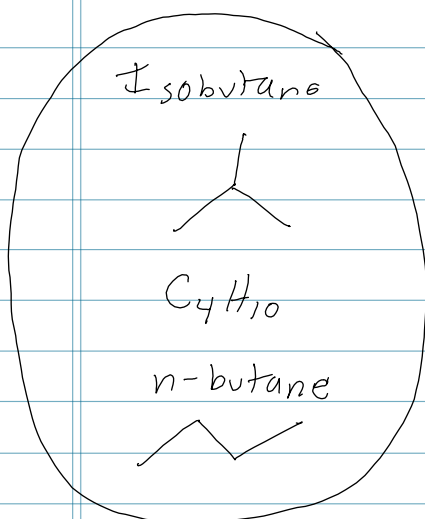
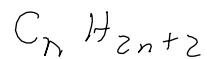
rotation in \sim kps



Alkanes

— p 2) Branched

also non-polar

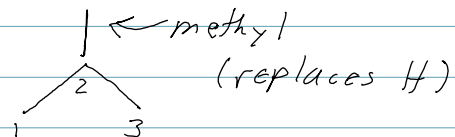


Constitutional Isomers: same chem. formula
different patterns of bonding

Systematic Nomenclature:

- 1) Longest C chain \Rightarrow base name
- 2) Number C's as 1, 2, ...
- 3) Substituent side chains named with -yl replacing -ane

Ex) Isobutane = 2-methylpropane



Isopentane = 2-methylbutane

Neopentane = 2,2-dimethylpropane

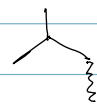
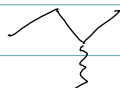
Special substituents (branchal)

isopropyl

tert-butyl

sec-butyl

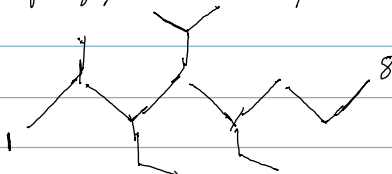
isobutyl



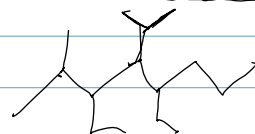
di = 2
tri = 3
tetra = 4
penta = 5
hexa = 6

Ex) 3,5-diethyl-4-isopropyl-2-methyloctane

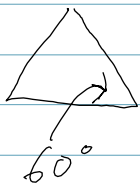
Wired Pencast drawing \rightarrow



regular drawing



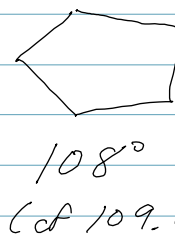
Cyclopropane



Cyclobutane



Cyclopentane

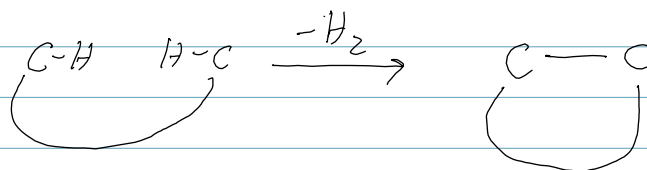


Strained angles
(less stable)

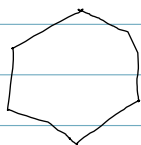
Acyclic alkanes: $C_n H_{2n+2}$

Cycloalkanes: $C_n H_{2n}$ (one ring) $C_n H_{2n-2}$ (2 rings)

ring forming



Cyclohexane ← non-planar



120°
No!

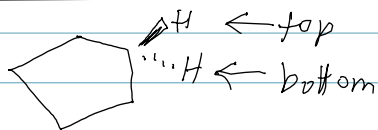


boat



chair
(most stable)

conformers
of each other
(cyclohexane)



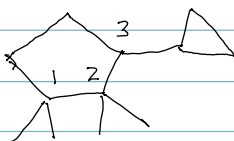
2 substituents: top + top = cis
top + bottom = trans

trans-4-(tert butyl)-3-ethyl-1,1-dimethylcyclohexane



$C_{14} H_{28}$

3-cyclopropyl-1,1,2,2-tetramethylcyclopentane



$C_{12} H_{22}$